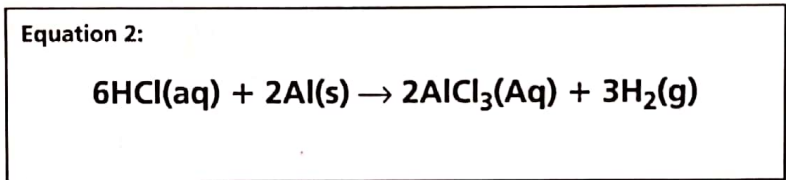
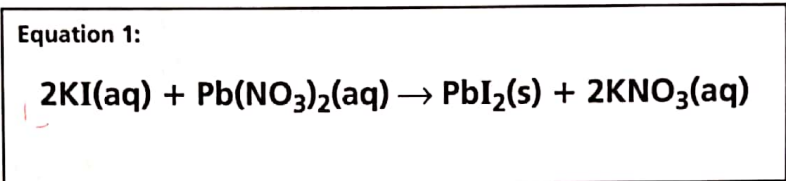
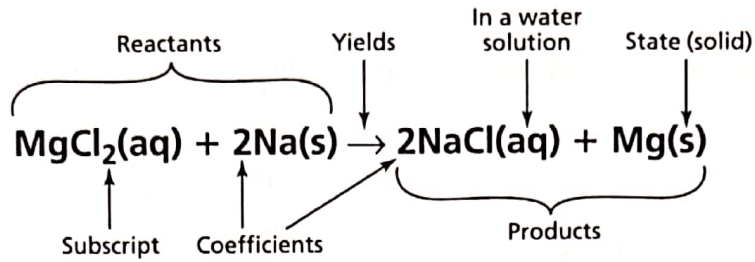


Parts of a Balanced Chemical Equation



1. What substances are reactants in located to (L) of arrow
 - a. Equation 1? KI, Pb(NO₃)₂ b. Equation 2? HCl, Al
2. What substances are products in located to (R) of arrow
 - a. Equation 1? PbI₂, KNO₃ b. Equation 2? AlCl₃, H₂
3. List the coefficients used in big #s in front (1 is NOT shown)
 - a. Equation 1? 2, 1, 1, 2 b. Equation 2? 6, 2, 2, 3
4. What substances are in aqueous solution in (aq)
 - a. Equation 1? KI, Pb(NO₃)₂, KNO₃ b. Equation 2? HCl, AlCl₃
5. What substance shown is a gas? (g) H₂
6. What is the state of PbI₂ in Equation 1? solid
7. What state is not represented in either equation? no liquid shown (l)

Counting Atoms

KEY
(HW)

Worksheet - independent practice (back side)

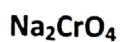
Count the atoms present in the different compounds by using the coefficients and subscripts.



Type of Atom	# of Atoms
K	2
C	1
O	3
Total	6



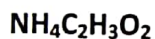
Type of Atom	# of Atoms
Ba	3
P	2
O	8
Total	13



Type of Atom	# of Atoms
Na	2
Cr	1
O	4
Total	7

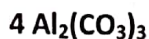


Type of Atom	# of Atoms
Ca	3
Cl	6
Total	9



Type of Atom	# of Atoms
N	1
H	7
C	2
O	2
Total	12

$4 + 3$



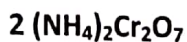
Type of Atom	# of Atoms
Al	8
C	12
O	36
Total	56

4×2
 $4 \times 1 \times 3$
 $4 \times 3 \times 3$



Type of Atoms	# of Atoms
Pb	1
N	2
O	6
Total	9

1×2
 3×2



Type of Atom	# of Atoms
N	4
H	16
Cr	4
O	14
Total	38

$2 \times 1 \times 2$
 $2 \times 4 \times 2$
 2×2
 2×7