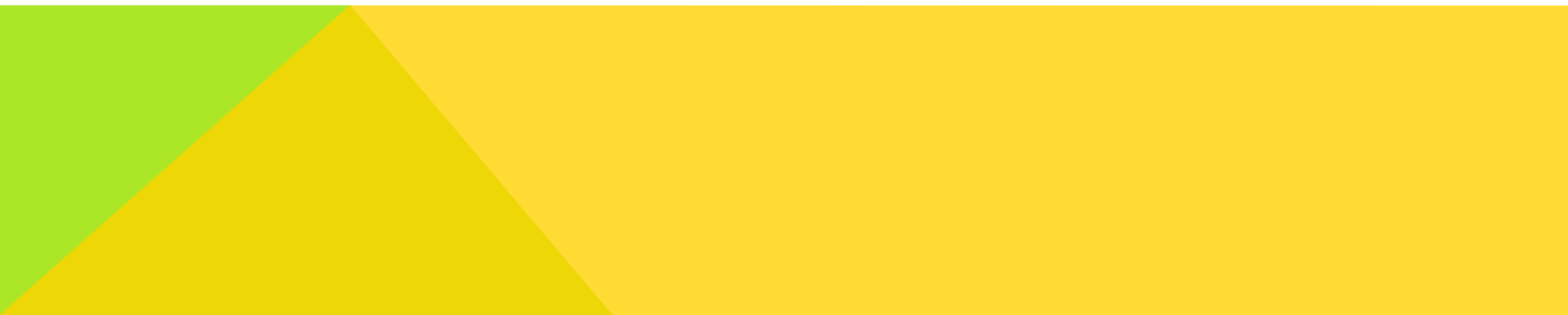


POLYATOMIC IONS

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Polyatomic Ion:



POLYATOMIC IONS

The importance of spelling...



POLYATOMIC IONS

Polyatomic Cations:

Ammonium NH_4^+

Hydronium H_3O^+

Polyatomic Anions:

Acetate $\text{C}_2\text{H}_3\text{O}_2^-$

Acetate CH_3COO^-

Bicarbonate HCO_3^-

Bromate BrO_3^-

Carbonate CO_3^{2-}

Chlorate ClO_3^-

Chlorite ClO_2^-

Chromate CrO_4^{2-}

Cyanide CN^-

Dichromate $\text{Cr}_2\text{O}_7^{2-}$

Hydroxide OH^-

Hypochlorite ClO^-

Nitrate NO_3^-

Nitrite NO_2^-

Oxalate $\text{C}_2\text{O}_4^{2-}$

Perchlorate ClO_4^-

Permanganate MnO_4^-

Peroxide O_2^{2-}

Phosphate PO_4^{3-}

Sulfate SO_4^{2-}

Sulfite SO_3^{2-}

POLYATOMIC IONS

Naming Rules: everything stays the same except...

- Polyatomic Ion

- Example: $\text{Na}_2\text{Cr}_2\text{O}_7$



POLYATOMIC IONS

a. KNO_3 _____

b. NH_4OH _____

c. $\text{Mg}(\text{C}_2\text{H}_3\text{O}_2)_2$ _____

d. $(\text{NH}_4)_2\text{S}$ _____

e. NaHCO_3 _____

f. CuSO_4 _____

POLYATOMIC IONS

Formula Rules:

- Use the cross over method
- If you need to put a subscript on the polyatomic ion,

- Ex: Aluminum Sulfate



POLYATOMIC IONS

Aluminum Sulfate



POLYATOMIC IONS

a. Sodium Oxalate

b. Lead II Nitrite

c. Ammonium Nitride

d. Potassium Permanganate

e. Calcium Phosphate

f. Titanium III Hydroxide

