

Name:

Period:

Naming Ionic & Covalent Compounds

Steps:

1. On the first line identify the compound as ionic (I) or covalent (C)- If it is covalent skip steps 2 and 3.
2. On the second line identify if the compound will need a roman numeral (Y) or not (N).
3. On the third line identify if the compound has a polyatomic ion (Y) or not (N).
4. Write the correct name.

_____ 1. $\text{Al}(\text{NO}_3)_3$ _____

_____ 2. FeCl_3 _____

_____ 3. CS_3 _____

_____ 4. TiO_2 _____

_____ 5. CaCO_3 _____

_____ 6. $\text{Cu}(\text{NO}_2)_2$ _____

_____ 7. $\text{Sn}(\text{CN})_4$ _____

_____ 8. $\text{Ba}_3(\text{PO}_4)_2$ _____

_____ 9. N_2O_4 _____

_____ 10. Ag_2SO_3 _____

_____ 11. NiSe_2 _____

_____ 12. P_3O_5 _____

_____ 13. $\text{Sn}_3(\text{PO}_4)_2$ _____

_____ 14. N_7Cl_3 _____

_____ 15. LiOH _____

Writing Formulas for Ionic & Covalent Compounds

Steps:

1. On the first line identify the compound as ionic (I) or covalent (C)- If it is covalent skip steps 2 and 3.
2. On the second line identify the charge of the metal.
3. On the third line identify the change of the nonmetal.
4. Write the correct formula.

_____ 16. Manganese (III) bromide _____

_____ 17. Calcium acetate _____

_____ 18. Sulfur dioxide _____

_____ 19. Tin (IV) sulfate _____

_____ 20. Zinc hydroxide _____

_____ 21. Lead (IV) nitride _____

_____ 22. Copper (II) chlorate _____

_____ 23. Carbon tetrachloride _____

_____ 24. Ammonium phosphate _____

_____ 25. Diphosphorus pentabromide _____

_____ 26. Barium cyanide _____

_____ 27. Cobalt (VI) chromate _____

_____ 28. Lithium nitrite _____

_____ 29. Trisulfur hexafluoride _____

_____ 30. Aluminum oxalate _____