

Name: _____

Period: _____

Writing Ionic Formulas from Names

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|-----------------------------|-------|--------------------------|-------|
| 1. Silver oxide | _____ | 12. Iron (II) fluoride | _____ |
| 2. Aluminum nitride | _____ | 13. Nickel (II) selenide | _____ |
| 3. Manganese (II) hydroxide | _____ | 14. Lithium oxide | _____ |
| 4. Ammonium carbonate | _____ | 15. Copper (I) carbonate | _____ |
| 5. Aluminum oxide | _____ | 16. Strontium carbonate | _____ |
| 6. Barium carbonate | _____ | 17. Potassium dichromate | _____ |
| 7. Calcium phosphate | _____ | 18. Manganese (II) oxide | _____ |
| 8. Magnesium sulfite | _____ | 19. Nickel (II) chloride | _____ |
| 9. Chromium (III) phosphide | _____ | 20. Lead (II) acetate | _____ |
| 10. Cobalt (III) nitrate | _____ | 21. Tin (IV) chloride | _____ |
| 11. Zinc iodide | _____ | | |

Writing Ionic Names from Formulas

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|---------------------------------------------------|-------|----------------------------------|-------|
| 22. $\text{Fe}_2(\text{SO}_4)_3$ | _____ | 33. Li_2SO_3 | _____ |
| 23. NiSO_4 | _____ | 34. FeC_2O_4 | _____ |
| 24. Li_3PO_4 | _____ | 35. AgCl | _____ |
| 25. PbSO_4 | _____ | 36. $(\text{NH}_4)_2\text{S}$ | _____ |
| 26. Mg_3N_2 | _____ | 37. $\text{Fe}(\text{NO}_2)_2$ | _____ |
| 27. $\text{Zn}(\text{C}_2\text{H}_3\text{O}_2)_2$ | _____ | 38. Bi_2O_3 | _____ |
| 28. $\text{Sr}(\text{OH})_2$ | _____ | 39. $\text{Mg}_3(\text{PO}_4)_2$ | _____ |
| 29. $\text{Pb}(\text{SO}_4)_2$ | _____ | 40. $\text{Ni}(\text{HCO}_3)_2$ | _____ |
| 30. NaOH | _____ | 41. $\text{Zn}(\text{OH})_2$ | _____ |
| 31. $\text{Ba}(\text{OH})_2$ | _____ | 42. AlPO_4 | _____ |
| 32. CaO | _____ | | |