*$2n^2$ determines how many electrons for each energy level

NAME_____

Bohr Model Practice

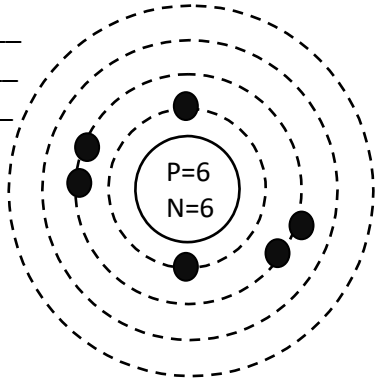
Protons and neutrons go in the nucleus (center). Electrons go in energy levels.

1st = 2e⁻ 2nd = 8e⁻ 3rd = 18e⁻ 4th = 32e⁻

Remember: The first energy level only holds up to 2 electrons. The second energy level holds up to 8 electrons.

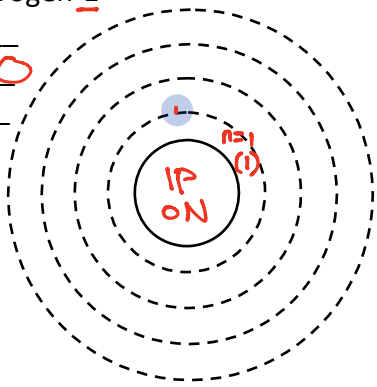
1. Carbon-12

P =____
N=____
E=____



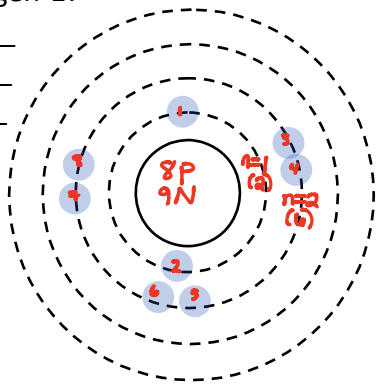
2. Hydrogen-1

P = 1
N=H=0
E= 1



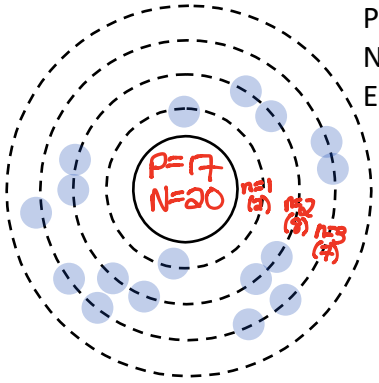
3. Oxygen-17

P = 8
N= 9
E= 8



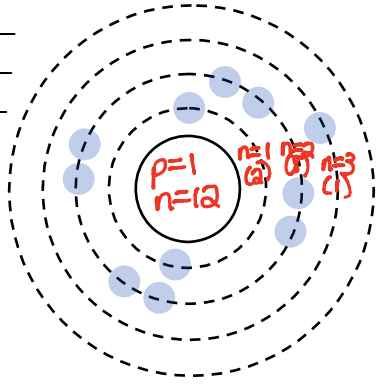
4. Chlorine-37

P = 17
N= 20
E= 17



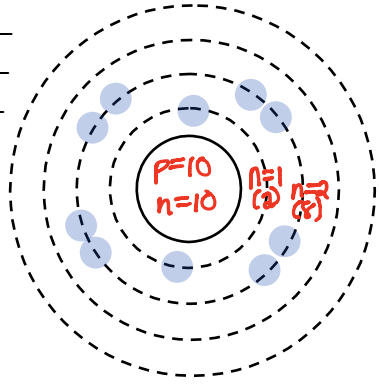
5. Sodium-23

P = 11
N= 12
E= 11



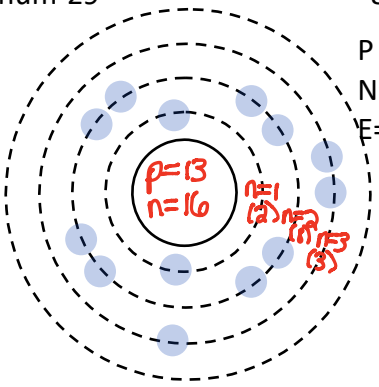
6. Neon-20

P = 10
N= 10
E= 10



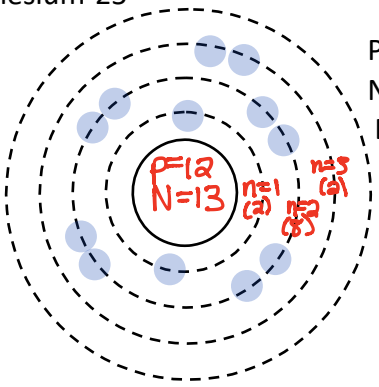
7. Aluminum-29

P = 13
N= 16
E= 13



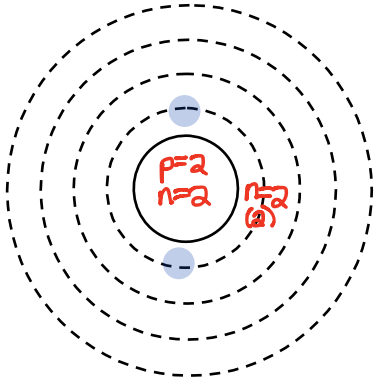
8. Magnesium-25

P = 12
N= 13
E= 12



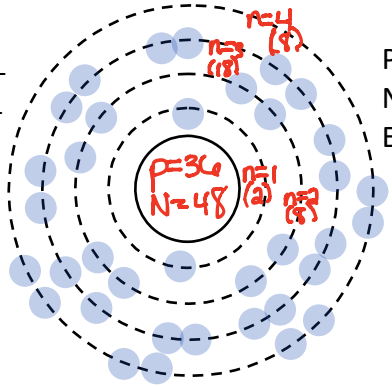
9. Helium-4

P = 2
N= 2
E= 2



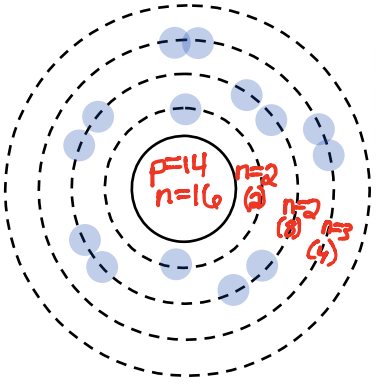
10. Krypton-84

P = 36
N= 48
E= 36



11. Silicon-30

P = 14
N= 16
E= 14



12. Beryllium-9

P = 4
N= 5
E= 4