

Periodic Trends Worksheet

Name KEY Per _____

1. Circle the element with the largest atomic radius and put a square around the element with the smallest atomic radius:

Cu K Ni Br

Explain why you made these choices. radius decreases across a period

2. Circle the element with the highest ionization energy and put a square around the element with the lowest ionization energy:

Cu K Ni Br

Explain why you made these choices. ionization energy increases across a period

3. Circle the element with the highest electronegativity and put a square around the element with the lowest electronegativity:

Cu K Ni Br

Explain why you made these choices. electronegativity increases across a period

4. Circle the element with the largest atomic radius and put a square around the element with the smallest atomic radius:

C Sn Si Ge

Explain why you made these choices. radius increases down a group

5. Circle the element with the highest ionization energy and put a square around the element with the lowest ionization energy:

C Sn Si Ge

Explain why you made these choices. ionization energy decreases down a group

6. Circle the element with the highest electronegativity and put a square around the element with the lowest electronegativity:

C Sn Si Ge

Explain why you made these choices. electronegativity decreases down a group

7. Circle the element with the largest atomic radius and put a square around the element with the smallest atomic radius:

O Al Ca

Explain why you made these choices. radius decreases from bottom left corner to top right corner

8. Circle the element with the highest electronegativity and put a square around the element with the lowest electronegativity:

Cl K Ba

Explain why you made these choices. electronegativity increases from bottom left corner to top right corner