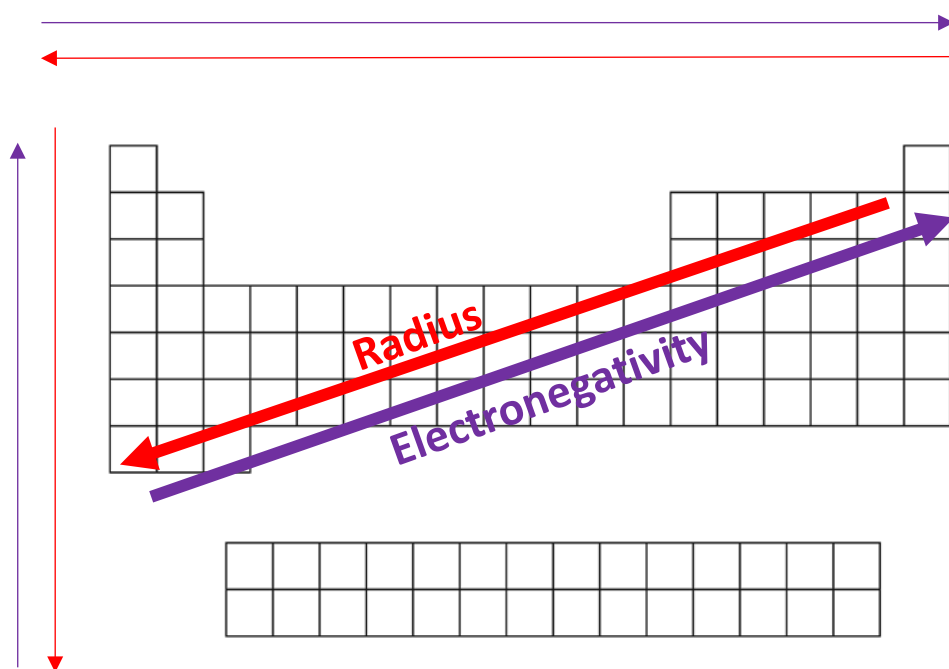


## Atomic Radius

- Definition: size of an **atom** (distance from the center of the nucleus to the last energy level)
- Groups: Radius increases going down because more and more energy levels are added
- Periods: Radius decreases going across a row because more protons are added which pull harder on the electrons

\*Remember your corners – F (smallest) and Fr (largest)\*



## Electronegativity

- How strongly an atom pulls (holds on to) on its electrons
- Groups: Electronegativity decreases going down a group because the atom gets larger and protons cannot pull on electrons as strongly
- Periods: Electronegativity increases going across a row because the atom gets smaller and protons pull on electrons stronger

\*Opposite trend of electronegativity – the larger the atom, the further away the electrons, which means less pull\*