

Notes: Isotopes

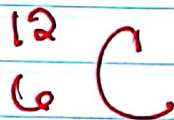
Different versions of an element

↳ mass (#) is different

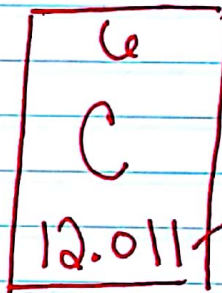
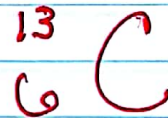
↳ comes from neutrons

You can estimate the most common mass # by looking at atomic mass

Carbon-12

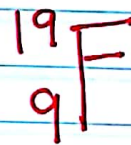
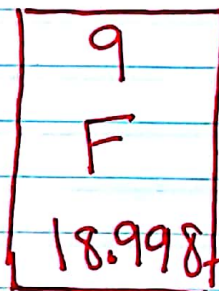


Carbon-13



→ closest to 12, so most common isotope must be Carbon-12 (with a mass # of 12)

ex:



→ closest to 19, so most common must be F-19

protons = 9
electrons = 9
neutrons = 10